**Chapter 3 Study Guide**

1. Mendeleev organized the periodic table how?
2. What is atomic number? Mass number? Atomic mass?
3. What is an isotope?
4. Draw and label an atom. What is between the nucleus and electrons?
5. From an element’s location in the periodic table, you can predict what?
6. How is the modern periodic table organized?
7. What is a group? How many are on the periodic table?
8. What is a period? How many are on the periodic table?
9. Classify elements on the periodic table:
	1. Alkali metals
	2. Alkaline earth metals
	3. Transition Metals
	4. Nonmetals
	5. Metalloids
	6. Halogens
	7. Nobel Gases
10. What are three differences between metals and nonmetals?
11. What are two ways that protons and neutrons are similar?
12. 3
13. How can an atom be electrically neutral when it contains particles that are charged?
14. Be able to identify the information in a square on the periodic table.



1. Explain why the atomic mass of an element is usually not given as a whole number even though each individual atom of the element has a whole number of protons and neutrons.
2. Why is iron a better building material than an alkali metal?